

# Units of Measurement and the Metric System

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## Objective:

To use metric system units of measurement in laboratory procedures and in problem solving, at the level of 85% proficiency for each student.

In order to achieve this objective, you will need to be able to:

1. Name the major metric units and their abbreviations for length, volume, and mass.
2. Define and use: kilo-, deci-, centi-, milli-, micro-, and nano-.
3. Recall how many inches in a m, cm in an inch, mL in a fluid ounce, liters in a quart, pounds in a kg, and mm in a cm.
4. Convert between the following: m, cm, mm,  $\mu\text{m}$ , nm.
5. Convert between the following: L, dL, mL,  $\mu\text{L}$ , nL.
6. Convert between the following: kg, g, mg,  $\mu\text{g}$ , ng.

## Materials

### Group Supplies:

Meter stick  
Millimeter ruler  
50 mL beaker  
10 mL graduated cylinder

### Lab Supplies:

Balance  
Water  
Paper clips  
Pencils

## Methods and Results

### ***Making conversions:***

1. Fill in the basic unit of metric measurement and their standard abbreviations:

	Name of Unit	Abbreviation
length	meter	m
volume (liquid)	liter	L
mass	gram	g
temperature	Celsius	C
food energy	calorie	cal

2. Fill in the prefixes and their abbreviations:

	Prefix	Abbreviation
One Million	mega	M
One Thousand	kilo	k
One Hundred*	hecto*	h
Ten*	deka*	da
One-Tenth	deci	d
One-Hundredth	centi	c
One-Thousandth	milli	m
One-Millionth	micro	$\mu$
One-Billionth	nano	n

\*Not commonly used

3. Write these numbers in decimal form and in scientific notation:

	Decimal	Scientific Notation
One Million	1000000.	$1 \times 10^6$
One Thousand	1000.	$1 \times 10^3$
One Hundred	100.	$1 \times 10^2$
Ten	10.	$1 \times 10^1$
One-Tenth	0.1	$1 \times 10^{-1}$
One-Hundredth	0.01	$1 \times 10^{-2}$
One-Thousandth	0.001	$1 \times 10^{-3}$
One-Millionth	0.000001	$1 \times 10^{-6}$
One-Billionth	0.000000001	$1 \times 10^{-9}$



- b. A = 1 mL; B = 2 fl oz; C = 3 L; D = 0.5 gallons; E = 0.75 pints; F = 2 tsp  
 60 mL      3000 mL      1893 ml              355 mL              ~10 mL

C	D	E	B	F	A
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- c. A = 2.3 lbs; B = 5 oz; C = 30 kg; D = 310 grams; E = 0.025 tons  
 1.0433 kg    0.1417 kg                      0.310 kg              22.68 kg

C	E	A	D	B
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### ***Additional Work with Metrics:***

1. Convert these numbers as indicated. Show your work including conversion factors and units (the first one has been done for you):

From:	Convert to:	show your work:	Answer	
0.45 L	mL	0.45 L x (1000 mL / L) = <table border="1" style="display: inline-table; vertical-align: middle;"><tr><td><math>0.45 \text{ L} \times \frac{1000 \text{ mL}}{\text{L}} =</math></td></tr></table>	$0.45 \text{ L} \times \frac{1000 \text{ mL}}{\text{L}} =$	450 mL
$0.45 \text{ L} \times \frac{1000 \text{ mL}}{\text{L}} =$				
1250 mL	L	1250 mL x (L / 1000mL) =	1.25 L	
0.065 mg	g	0.065 mg x (g / 1000 mg) =	0.000065 g	
3.7 km	m	3.7 km x (1000 m / km) =	3700 m	
120 cm	km	120 cm x (km / 100,000 cm) =	0.0012 km	
3.6 kg	g	3.6 kg x (1000 g / kg) =	3600 g	
670 cm	m	670 cm x (m / 100 cm) =	6.7 m	
1250 g	kg	1250 g x (kg / 1000 g) =	1.25 kg	
0.15 L	mL	0.15 L x (1000 mL / L) =	150 mL	
120 mm	cm	120 mm x (cm / 10 mm) =	12 cm	
627 L	mL	627 L x (1000 mL / L) =	627,000 mL	

2. You have to give your dog medicine at a dose rate of 1.5 mg of medicine per kg of the dog's weight. The dog weighs 50 lbs. How much medicine should you give him? (Show your work.)  $50 \text{ lbs} / 2.2046 \text{ lbs} / \text{kg} = 22.68 \text{ kg} \times 1.5 \text{ mg} / \text{kg} = 34.02 \text{ mg}$

Answer: 34.02 mg

3. You have a fever and your temperature is 102 degrees Fahrenheit.  
 a. What is your temperature in degrees Celsius?  $102 - 32 = 70 \times 5/9 = 38.89$   
 b. When your temperature returns to normal (98.6 degrees F) what is your temperature in Celsius? (Show your work.)  $98.6 - 32 = 66.6 \times 5/9 = 37$

a. Answer: 38.89 C

b. Answer: 37 C

4. You have a friend who is 74 inches tall. How tall is your friend: (Show your work)

a. in centimeters?

Answer: 187.96

$74 \text{ in} \times 2.54 \text{ cm} / \text{in}$

b. in millimeters?

Answer: 1879.6 mm

$74 \text{ in} \times 25.4 \text{ mm} / \text{in}$

c. in meters?

Answer: 1.8796 m

$74 \text{ in} \times 0.0254 \text{ m} / \text{in}$

**Table of Metric and English Equivalents and English Conversions for Common Measurements.**

Quantity	Metric Unit	Symbol	Metric equivalent	English equivalent	English conversion
Length	kilometer	km	1000 m		
	<b>meter</b>	<b>m</b>	1 m	1.0936 yds	1 yd = 0.9144 m
	decimeter	dm	0.1 m		
	centimeter	cm	0.01 m	0.3937 in	1 in = 2.54 cm
	millimeter	mm	0.001 m		
	micrometer	μm	0.000001 m		
	nanometer	nm	0.000000001 m		
Area	<b>square meter</b>	<b>m<sup>2</sup></b>		10.7639 ft <sup>2</sup>	1 ft <sup>2</sup> = 0.0929 m <sup>2</sup>
	square centimeter	cm <sup>2</sup>	0.0001 m <sup>2</sup>		
Volume	cubic meter	m <sup>3</sup>	1000 L		
	<b>liter</b>	<b>L</b>	1 L	1.0567 qt	1 qt = 0.94635 L
	deciliter	dL	0.1L		
	cubic centimeter	cm <sup>3</sup>	0.001 L		
	milliliter	mL	0.001 L	0.0338 fl oz	1 fl oz = 29.5735 mL
					1 tsp = 4.9289 mL
	microliter	μL	0.000001 L		
nanoliter	nL	0.000000001 L			
Mass	kilogram	kg	1000 g	2.2046 lb	1 lb = 0.4536 kg
	<b>gram</b>	<b>g</b>	1.g	0.0353 oz	1 oz = 28.3495 g
	<b>gram</b>	<b>g</b>	1 mL water @ 4 °C		
	milligram	mg	0.001 g		
	microgram	μg	0.000001 g		
	nanogram	ng	0.000000001 g		
	picogram	pg	0.000000000001 g		
Energy	<b>celsius</b>	°C		(5/9) x (F - 32)	F = ((9/5) x C) + 32
	<b>calorie</b>	cal*	4.186 joules		
	Calorie**	Cal	4.186 kilojoules		

\*amount of heat energy required to raise 1 mL of water by 1 °C

\*\*Also known as a kilocalorie (kcal)

## Discussion:

1. Explain the importance of the metric system in medicine.
2. Explain the meaning of the prefixes kilo(k)-, deci(d)-, centi(c)-, milli(m)-, micro( $\mu$ )-, and nano(n)-.
3. Explain how to convert a metric unit with one prefix into the same metric unit with another prefix. (For example, explain how to convert 1 dL to \_\_mL.)